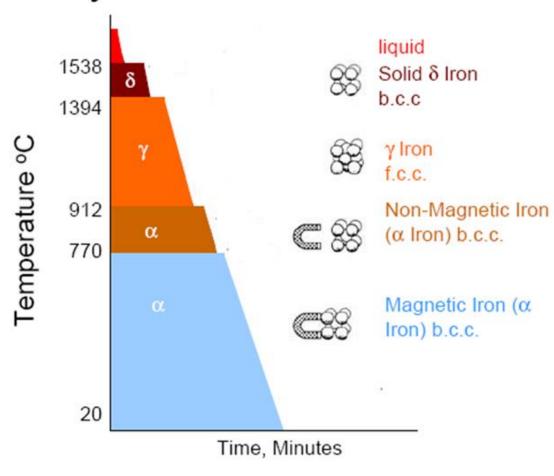
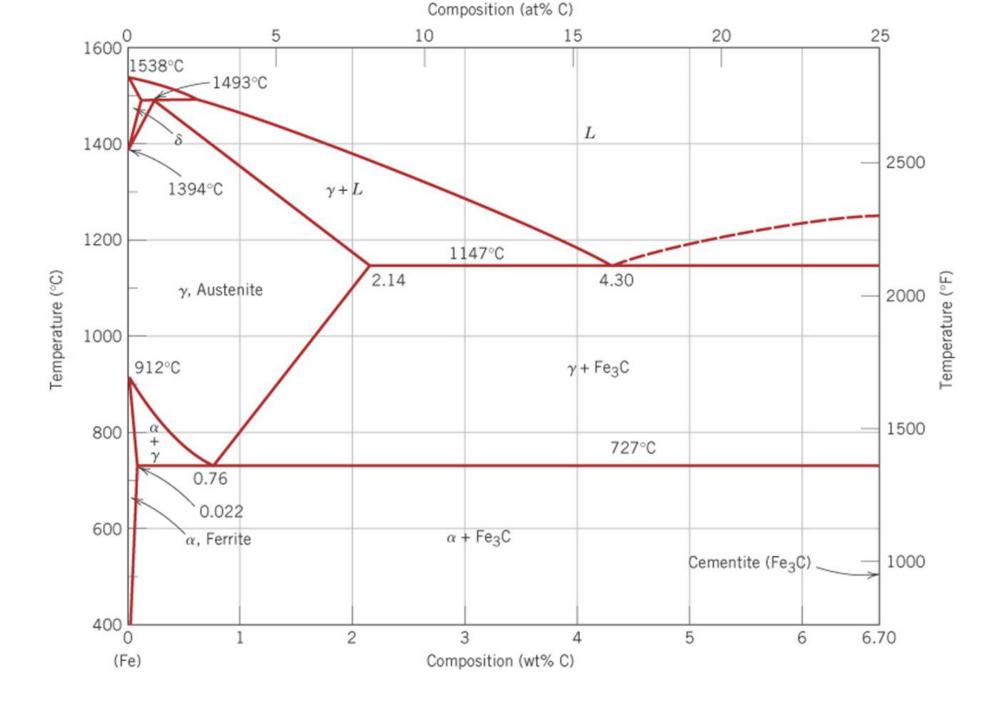
آزمایشگاه متالوگرافی

دیاگرام تعادلی آهن-کربن

- Pure iron when heated experiences 2 changes in crystal structure before it melts.
- At room temperature the stable form, ferrite (α iron) has a BCC crystal structure.
- Ferrite experiences a polymorphic transformation to FCC austenite (γ iron) at 912 °C (1674 °F).
- At 1394°C (2541°F) austenite reverts back to BCC phase δ ferrite and melts at 1538 °C (2800 °F).
- Iron carbide (cementite or Fe₃C) an intermediate compound is formed at 6.7 wt% C.
- Typically, all steels and cast irons have carbon contents less than 6.7 wt% C.
- Carbon is an interstitial impurity in iron and forms a solid solution with the α , γ , δ phases.

Crystal structures of iron





4 Solid Phases

α-ferrite

- solid solution of carbon in a iron,
- BCC structure
- carbon only slightly soluble in the matrix
 - maximum solubility of 0.02%C at 723°C to about 0.008%C at room temperature.

Austenite (γ)

- solid solution of carbon in γ -iron
- FCC structure: can accommodate more carbon than ferrite
 - maximum of 2.08%C at 1148°C, decreases to 0.8%C at 723°C
 - difference in C solid solubility between γ and α is the basis for hardening of most steels.

δ -ferrite

- solid solution of carbon in δ–iron
- BCC crystal structure
 - maximum solubility of ferrite being 0.09%C at 1495°C

Cementite (Fe₃C)

- intermetallic Fe-C compound
 - Fe₃C: 6.67%C and 93.3%Fe.
- orthorhombic crystal structure: hard and brittle

Iron carbide (Cementite or Fe₃C)

- Forms when the solubility limit of carbon in α ferrite is exceeded at temperatures below 727 °C.
- Mechanically, cementite is very hard and brittle.
- For ferrous alloys there are 3 basic types, based on carbon content:
 - ☐ Iron (ferrite phase): <0.008 wt% C room temp
 - \square Steel (α + Fe₃C phase): 0.008 to 2.14 wt% C
 - ☐ Cast iron: 2.14 to 6.70 wt% C